**Inglés Básico para Química**

**Tipos de texto**: instructivo, expositivo, descriptivo, narrativo y argumentativo.

**Text 1:**

Lavoisier was born into a very affluent family in Paris. When he was five, his mother passed away, and she left him a large fortune. In the mid 1700’s he attended the Collège Mazarin and studied astronomy, botany, chemistry and mathematics. He obtained a law license in 1764 before embracing a career in science. When he was 25, he became a member France’s elite scientific organization, the French Academy of Sciences. His first science publication occurred in 1764. He worked on France’s first geological map in 1769.

<http://famouschemists.org/antoine-lavoisier/>

**Text 2:**

[Benzoic acid](http://www.rsc.org/learn-chemistry/wiki/Substance%3ABenzoic_acid) in its neutral form is of fairly moderate polarity (as is [benzil](http://www.rsc.org/learn-chemistry/wiki/Substance%3ABenzil%22%20%5Co%20%22Substance%3ABenzil)), and it is soluble in medium polarity solvents such as [ethyl acetate](http://www.rsc.org/learn-chemistry/wiki/Substance%3AEthyl_acetate%22%20%5Co%20%22Substance%3AEthyl%20acetate) (as is benzil). Ethyl acetate is a solvent that has low toxicity, and it does not mix with water. If we take an ethyl acetate solution of our mixture and add a base such as sodium bicarbonate, benzoic acid will form its sodium salt (i.e., its conjugate base), and this ionic compound will have markedly different solubility properties from the acid. For example, it will be soluble in cold water, unlike both benzil and neutral benzoic acid. This sodium salt can be removed as an aqueous solution, from which the uncharged benzoic acid can be recovered simply by acidification.

[http://www.rsc.org/learn-chemistry/wiki/Expt:Separation\_of\_an\_Acid\_from\_a\_Neutral\_by\_Base\_Extraction](http://www.rsc.org/learn-chemistry/wiki/Expt%3ASeparation_of_an_Acid_from_a_Neutral_by_Base_Extraction)

**Text 3:**

Dissolve about 1.00 g of the benzoic acid/benzil (1:1) mixture in 10 mL of ethyl acetate in a 50 mL Erlenmeyer flask, and label this flask 1. Then add 5 mL of 1.0 M sodium bicarbonate to the flask (note what happens), mix well, then pour off the mixture into a separatory funnel. Mix the contents thoroughly by shaking vigorously for about 2-3 min with periodic venting (your instructor will demonstrate the correct technique for this). Allow the layers to separate completely, then drain off the lower layer into a second 50 mL Erlenmeyer flask (labeled flask 2). Add another 2 mL of sodium bicarbonate solution to the separatory funnel, mix the contents as before, and drain off the lower layer into flask 2. Exactly what chemical compound is present in flask 2? Drain off the upper layer into a 50 mL round bottomed flask, and label this "flask 3". Return the contents of flask 2 into the separatory funnel, then add 2 mL of ethyl acetate and mix it thoroughly. Drain off the lower layer back into flask 2, then discard the upper ethyl acetate layer into the designated waste. This is called backwashing and serves to remove any organic material that might contaminate the contents of flask 2.

[http://www.rsc.org/learn-chemistry/wiki/Expt:Separation\_of\_an\_Acid\_from\_a\_Neutral\_by\_Base\_Extraction](http://www.rsc.org/learn-chemistry/wiki/Expt%3ASeparation_of_an_Acid_from_a_Neutral_by_Base_Extraction)

**Text4:**

What we need is to bring thinking about experimental chemistry back into our introductory organic chemistry labs, in projects and experiments that involve asking questions and seeking answers using well designed procedures. I am convinced that this question-driven, guided-inquiry approach can bring back the life into organic chemistry teaching laboratories and can let students participate effectively in the process of our science.

<http://pubs.acs.org/doi/pdf/10.1021/ed081p1083>